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Statistical Thermodynamics Paperback – December 1, 1984 by Donald A. McQuarrie (Author)

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Problem 3-4. Fall 2003. We have to derive the thermodynamic properties of an ideal monatomic gas from the following: $\Omega = \frac{1}{N!} \left(\frac{2\pi m kT}{h^2} \right)^{3N/2} \frac{1}{V^N} e^{-\beta U}$ the partition function for the grand canonical ensemble, where T , V , are fixed.

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Given that entropy is a measure of the dispersal of energy in a system, the more chaotic a system is, the greater the dispersal of energy, the greater the entropy will be. Ludwig Boltzmann (1844 – 1906) (O'Connor & Robertson, 1998) understood this concept well, and used it as a statistical approach to calculating entropy.

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Don McQuarrie's research interests were in theoretical chemistry, specifically in statistical thermodynamics of electrolyte systems and membranes, stochastic differential equations, and nonlinear dynamical systems. He authored or co-authored 10 textbooks and 95 scientific papers.

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