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compounds

composed of two

different

elements;

(compounds of 2

elements

between metal

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***(positive, s-block
and p-block) and
nonmetal
(negative); metal
is always listed
first; change
nonmetal to -ide
ending)***

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Chapter Test

Assessment

Chapter Test B

Teacher Notes

and Answers 5

The Periodic Law

TEST B 1. a 2. c

3. d 4. d 5. a 6. a

7. c 8. a 9.

lanthanides 10. 2

11. fourth 12.

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transition

elements 13. 32

14. valence

electrons 15.

electron affinity

16.

electronegativity

17. ionization

energy 18. 3 s²

3p⁴ 19. atomic

radius 20. ion 21.

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LESSON 7 1**

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The Chemical

Formulas and

Chemical

Compounds

chapter of this

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**Chemistry
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students learn
the essential
lessons
associated with
chemical
compounds and
...**

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Instructions:

Choose your

answers to the

questions and

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see the next set

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come back to

them later with

the yellow "Go

To First Skipped

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Question"
Answer Key

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Chapter 7 test.
STUDY. PLAY.
The basis of an
ionic bond is
the absorption**

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Answer Key

**of water into
their solid
structures.**

**Elements tend to
gain or lose
electrons in
order to acquire
the electron
configuration of
a noble gas.**

**Ions that are
formed from
more than one
atom are called** —

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Answer Key

**Assessment
Chapter Test B
Holt McDougal
Modern
Chemistry 1
Chapter Test
Assessment
Chapter Test A
Chapter: The
Periodic Law Use
the periodic
table below to
answer the**

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Answer Key
**questions in this
Chapter Test. In
the space
provided, write
the letter of the
term or phrase
that best
completes each
statement or
best answers
each question.**

_____ 1.

Assessment

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Answer Key

Chapter Test A
Modern

Chemistry 71

Chapter Test

Name Class Date

**Chapter Test B,
continued 17. A**

substance

combines with

oxygen,

**releasing a large
amount of**

energy as heat

and light, in a(n)

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. 18. The decomposition of a substance by an electric current is called .

19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

Assessment

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**Chapter Test B -
Answer Key**
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Chemistry

Chapter

Resources.

Chapter One

Matter &

Change.

Objectives:

Chapter 1

Honors: ... CH 1

Reading

Assignment

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Modern

Chemistry. CH 1

Vocabulary-New.

CH 1 Mixed

Questions. ... CH

7 Pre Test .

White Board

Practice. CH 7

Ionic & Covalent

Nomenclature.

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Answer Key
**Chapter 7:
Chemical ...**

Yay, this practice test will determine if you're ready to take the real thing. Questions are taken right out of the book/packet and review lotsa vocab.

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CHAPTER 7

REVIEW.

Chemical

Formulas and

Chemical

Compounds.

SECTION 1.

SHORT ANSWER

Answer the

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**following
questions in the
space provided.**

**1. c In a Stock
system name
such as iron(III)
sulfate, the
Roman numeral
tells us. (a) how
many atoms of
Fe are in one
formula unit. (b)
how many
sulfate ions can**

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Answer Key
**be attached to
the iron atom.**

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Modern
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Chapter Test
Name Class Date
Chapter Test A,**

continued _____ 7.

The melting

points of ionic

compounds are

higher than the

melting points of

molecular

compounds

because a. ionic

substances tend

to vaporize at

room

temperature. b.

ionic substances

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are brittle. c.

Answer Key

**attractive forces
between ions are
greater than the
attractive ...**

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chapter 7

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Chemical ...

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Formulas and

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Compounds

Chemical Names

& Formulas,

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**Oxidation
Numbers, Using
Chemical
Formulas, &
Determining
Chemical
Formulas Terms
STUDY**

**7 Chemical
Formulas and
Chemical
Compounds
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Chapter Test

Name Class Date

**Chapter Test B,
continued PART**

**III Match the
letter on the
diagram with the
term for that
point or region.**

18. Critical point

19. Vapor 20.

**Normal freezing
point 21. Triple**

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Chapter 7 Test
Answer Key
point 22. Solid
23. Normal

boiling point 24.
Liquid 25.

Critical pressure

PART IV Write

the answers to
the following
questions in the
space provided.
26.

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Answer Key

**Chapter 7
Chemical
Formulas and
Chemical ...
Rules 1-7 apply
covalently
bonded atoms;
Oxidation
numbers can
also be assigned
to atoms in ionic
compounds.
Oxidation Rule 9
Monoatomic ion**

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**oxidation
number equal to
ion charge Na^+ ,
 Ca^{2+} , Cl^-**

Assessment

Chapter Test A

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