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Sections 1-3

Introduction to

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Calculations Limiting

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Yield 2 Definitions.

Composition

stoichiometry deals

with the mass

relationships of

elements in

compounds. Reaction

stoichiometry involves

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the mass  
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answer on the line to  
the left. Show all your  
work in the space

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provided. 1. \_\_\_\_\_

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2.

6.0 mol of  $N_2$  are mixed with 12.0 mol of H

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### **Stoichiometry**

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ANSWER Answer the following questions in the space provided. 1.

Given the following equation:  $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$  \_\_\_\_\_ a.

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301 Section Review



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page 301 Using

Conversion Factors

Page 306 #1 Page

306 #2 Chapter 9 quiz

#1- mole to mole

calculations  $6 \text{ NaBr} +$

$\text{Mg}_3(\text{PO}_4)_2 \rightarrow$

$\text{Na}_3\text{PO}_4 + 3 \text{ MgBr}_2$

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& #38 Answers . 38.

To ensure that all  
magnesium is  
converted to  $\text{MgO}$ , I  
would use pure  
oxygen, not air, to  
carry out the reaction,  
because Mg could  
react with N<sub>2</sub> in air to

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ANSWER Answer the  
following questions in

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the space provided. 1.

Given the following

equation:  $C_3H_4(g) +$

$xO_2(g) \rightarrow 3CO_2(g) +$

$2H_2O(g)$  4 a. What is

the value of the

coefficient x in this

equation? 40.07 g/mol

b. What is the molar

mass of  $C_3H_4$ ? 2

mol  $O_2$ :1 mol  $H_2O$  c.

What is the mole ratio

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