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b. the same as covalent bonding. c. a combination of ionic and covalent bonding. different from ionic or covalent bonding. 2. the valence electrons in a metallic bond move freely throughout the network of metal atoms. b. are held tightly by the most positively charged atom. c. are shared equally between two metal atoms.

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Assessment Chapter Test A

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Explain your answer. The three oxygen atoms have oxidation states of 6 total, and because the algebraic sum of the oxidation states in a neutral compound must be zero, the two nitrogen atoms must have oxidation states of 6 total, therefore 3 each.

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### Assessment Chapter Test B

Modern Chemistry 88 Chapter Test Chapter: States of Matter PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. \_\_\_\_ 1. According to the kinetic-molecular theory, particles of an ideal gas a. attract each other but do not collide. b. repel each other and collide.

### 6 Chemical Bonding - Effingham County Schools / Overview

Modern Chemistry 86 Chapter Test Name Class Date Chapter Test A, continued \_\_\_\_ 13. The separation process of paper chromatography can be explained by a. vaporization of the ink. b. water vapor pressure. c. capillary action. d. pull of gravity. \_\_\_\_ 14. When there is a small decrease in temperature, the average kinetic energy of the particles of a liquid

### 7 Chemical Formulas and Chemical Compounds

Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. b In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds. 2. a The fact that metals are malleable and ionic crystals are brittle is best

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### Modern Chemistry Test B Answer

Assessment Chapter Test B - clarkchargers.org. Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19.

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Holt McDougal Modern Chemistry Chapter Test Assessment Chapter Test B Teacher Notes and Answers 5 The Periodic Law TEST B 1. a 2. c 3. d 4. d 5. a 6. a 7. c 8. a 9. lanthanides 10. 2 11. fourth 12. transition elements 13. 32 14. valence electrons 15. electron affinity 16. electronegativity 17. ionization energy 18. 3 s<sup>2</sup> 3p<sup>4</sup> 19. atomic radius 20. ion 21.

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b. Br<sup>-</sup>, bromide ion. c. Br<sup>+</sup>, bromium ion. d. Br<sup>-</sup>, bromium ion. 2. The platinum(II) ion and the platinum(IV) ion a. are anions. b. are polyatomic ions. c. have charges of 2+ and 4+, respectively. d. have charges of 1+ and 3+, respectively. 3. is the

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name of the compound made of zinc ions, Zn and fluoride ions, F — ? a. zinc difluoride b. zinc fluorate

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a. Meyer c. Stas b. Ramsay d. Moseley \_\_\_\_ 13. The most useful source of general information about the elements for anyone associated with chemistry is a. a. calculator. c. periodic table. b. table of metric equivalents. d. table of isotopes. \_\_\_\_ 14. The periodic table. a. permits the properties of an element to be predicted before the element ...

Modern Chemistry Chapter 3 Test B Answers

Modern Chemistry: Chapter Tests with Answer Key, 2006, Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, Harcourt School Publishers, 2006 ... Michigan Holt Chemistry and Modern Chemistry Test Preparation Workbook: Help for the MME , Holt, Holt Rinehart & Winston, Jun 1, 2006, Juvenile Nonfiction, 126 pages. .

Assessment Chapter Test B - Ed W. Clark High School

Holt McDougal Modern Chemistry 3 Chapter Test Chapter Test B, continued 16

Modern chemistry chapter 3 test b answers. The measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called \_\_\_\_\_. 17. The energy required to remove one electron from an atom is called its \_\_\_\_\_. 18.

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds MIXED REVIEW  
SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO<sub>3</sub> a. copper(II) carbonate Na<sub>2</sub>SO<sub>3</sub> b. sodium sulfite (NH<sub>4</sub>)<sub>3</sub>PO<sub>4</sub> c. ammonium phosphate SnS<sub>2</sub> d. tin(IV) sulfide HNO<sub>2</sub> e. nitrous acid 2.

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Name CHAPTER 3 TEST continued Date Class FILL IN THE BLANK Write the correct term (or terms) in the space provided. 9. If a particular compound is composed of elements A and B, the ratio of the mass of B to the mass of A will always be the same.

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