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Flashcards and ... - Quizlet
Attraction between metal atoms and the surrounding sea of elec...
Chemical Bond Forces that hold groups of atoms together and make them functi...
Ionic Bonding
Chemical bonding that results from the electrical attraction b...
A chemical formula that shows the number and kinds of atoms in...
The energy required to break the bonds in one mole of a chemic....

Assessment Chapter Test B

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Modern Chemistry 88 Chapter
Test Chapter: States of Matter
PART I In the space provided,
write the letter of the term or
phrase that best completes each
statement or best answers each
question. _____ 1. According to the
kinetic-molecular theory, particles
of an ideal gas a. attract each other
but do not collide. b. repel each
other and collide.

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Explain your answer. The three
oxygen atoms have oxidation
states of 6 total, and because the
algebraic. sum of the oxidation
states in a neutral compound must
be zero, the two nitrogen. atoms
must have oxidation states of 6
total, therefore 3 each.

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CHAPTER 7 REVIEW Chemical
Formulas and Chemical
Compounds MIXED REVIEW

SHORT ANSWER Answer the following questions in the space provided. 1. Write formulas for the following compounds: CuCO₃ a. copper(II) carbonate Na₂SO₃ b. sodium sulfite (NH₄)₃PO₄ c. ammonium phosphate SnS₂ d. tin(IV) sulfide HNO₂ e. nitrous acid 2.

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B Answers

Modern Chemistry: Chapter Tests
with Answer Key, 2006, Holt
Rinehart & Winston, Holt, Rinehart
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Chemistry and Modern Chemistry
Test Preparation Workbook: Help
for the MME , Holt, Holt Rinehart
& Winston, Jun 1, 2006, Juvenile
Nonfiction, 126 pages. .

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b. the same as covalent bonding. c.
a combination of ionic and covalent
bonding. different from ionic or
covalent bonding. 2. the valence
electrons in a metallic bond move
freely throughout the network of
metal atoms. b. are held tightly by
the most positively charged atom.
c. are shared equally between two
metal atoms.

Assessment Chapter Test A
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Chapter 8

Answer Key

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TEST B 1. a 2. c 3. d 4. d 5. a 6. a
7. c 8. a 9. lanthanides 10. 2 11.
fourth 12. transition elements 13.
32 14. valence electrons 15.
electron affinity 16.
electronegativity 17. ionization
energy 18. $3s^2 3p^4$ 19. atomic
radius 20. ion 21.

New Page 1 [srvhs.org]

Name CHAPTER 3 TEST

continued Date Class FILL IN THE
BLANK Write the correct term (or
terms) in the space provided. 9, If
a particular compound is composed
of elements A and B, the ratio of
the mass of B to the mass of A will

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always be the same.

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a. Meyer c. Stas b. Ramsay d.
Moseley ____ 13. The most useful
source of general information
about the elements for anyone
associated with chemistry is a. a.
calculator. c. periodic table. b.
table of metric equivalents. d. table
of isotopes. ____ 14. The periodic
table. a. permits the properties of
an element to be predicted before
the element ...

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Assessment Chapter Test B -
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Chemistry 71 Chapter Test Name
Class Date Chapter Test B,
continued 17. A substance

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combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19.

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3 Chapter Test Chapter Test B,

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Chapter 8

measure of the ability of an atom in a chemical compound to attract electrons from another atom in the compound is called _____. 17. The energy required to remove one electron from an atom is called its _____. 18.

quiz chapter 6 modern chemistry
Flashcards and Study Sets ...

b. Br^- , bromide ion. c. Br^+ , bromium ion. d. Br^- , bromium ion.

2. The platinum(II) ion and the platinum(IV) ion a. are anions. b. are polyatomic ions. c. have charges of $2+$ and $4+$, respectively. d. have charges of $1+$ and $3+$, respectively. 3. is the name of the compound made of zinc ions, Zn and fluoride ions, F^- ? a. zinc difluoride b. zinc fluorate

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6 Chemical Bonding - Effingham
County Schools / Overview
Modern Chemistry 86 Chapter
Test Name Class Date Chapter
Test A, continued ____ 13. The
separation process of paper
chromatography can be explained
by a. vaporization of the ink. b.
water vapor pressure. c. capillary
action. d. pull of gravity. ____ 14.
When there is a small decrease in
temperature, the average kinetic
energy of the particles of a liquid

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m

CH 1 Reading Assignment Modern
Chemistry. CH 1 Vocabulary-New.
CH 1 Mixed Questions. CH 1
Matter & Energy Vocabulary-New.
CH 1 Classifying Matter-New. CH
1 Classification of Matter ... CH 11

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Pre Test Answers. CH 11 Section
Review 11-1 - 11-3 . CH 11 Gas
Stoichiometry 1 . CH 11 Gas
Stoichiometry 2. CH 11 Boyles
Law . CH 11 Charles Law ...

7 Chemical Formulas and Chemical
Compounds

Chemical Bonding SECTION 4

SHORT ANSWER Answer the
following questions in the space
provided. 1. b In metals, the

valence electrons are considered
to be (a) attached to particular
positive ions. (c) immobile. (b)
shared by all surrounding atoms.

(d) involved in covalent bonds. 2. a
The fact that metals are malleable
and ionic crystals are brittle is
best

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