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Arrhenius theory - Arrhenius acid and base, examples

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Revision Notes on Coordination Compounds. Ligands: an ion or molecule capable of donating a pair of electrons to the central atom via a donor atom. Unidentate ligands: Ligands with only one donor atom, e.g. NH₃, Cl⁻, F⁻ etc. Bidentate ligands: Ligands with two donor atoms, e.g. ethylenediamine, C₂O₄²⁻ (oxalate ion) etc. Tridentate ligands: Ligands which have three donor atoms per ligand, e ...

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Chemistry Olympiad past papers. ... (H₂O) 6] 2+ (aq) (pale blue) ... H₂O, NH₃ and Cl⁻ can act as monodentate ligands. The ligands NH₃ and H₂O are similar in size and are uncharged. Exchange of the ligands NH₃ and H₂O occurs without change of co-ordination number (eg Co²⁺ and Cu²⁺). ...

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Arrhenius Base. An Arrhenius base is a substrate that increases the concentration of hydroxide ions in the aqueous solution. The example for Arrhenius base is highly soluble sodium hydroxide compound in water, which dissociates to give sodium ion and hydroxide ion.

Revision Notes on Co-Ordination Compounds | askITians

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