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$5 \text{ C}_6\text{H}_5\text{COOH}(\text{s}) + 7.5 \text{ O}_2(\text{g}) \rightarrow 7 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l})$  and  $n$  is equal to 7 minus 7.5, or minus 0.5. Hence for the combustion of one mole of benzoic acid at 25°C  $\Delta H = -3227 \text{ kJ}$   
 $8.31 \text{ J K}^{-1} \text{ mol}^{-1} \times 298 \text{ K} = -2480 \text{ J}$

APPARATUS AND CHEMICALS

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solution. We will assume the solution is ideal, Raoult's law is

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$\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2 \cdot x\text{H}_2\text{O}$  by titration against a standard solution of potassium manganate (VII) (permanganate). Introduction This experiment provides useful experience in preparation for a practical examination. The directions are similar to those that would be given by an examination board.

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