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LHS AP Chemistry
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Solutions. $[H^+ O^-] =$

1.38×10^{-4} ; $pH=3.86$

Using this chart and
the explanation I
provide you with, you
will conduct a short
experiment in search
of which weak-acid-
conjugate pair is the
right buffer for you (or
at least your assigned

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challenge)! Ammonia is a weak base. The pK value for this reaction is 4.7. $pK = 14 - 4.7 = 9.3$.

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buffer using 55 mL of
0.5 M acetic acid and
45 mL of 0.5 M
sodium acetate.

Record the initial pH
and then add 10 mL
of 0.2 HCL to 25 mL
of the buffer solution
then record the pH.
Repeat using 0.2 M
NaOH. Record results
in appropriate data
tables and graphs.

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students attempt to
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solution effective in a
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properties of buffers,
as does the essential
function of the
respiratory system,
breathing, which must
be regulated within a
very narrow pH range.

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provide an essential acid-base balancing act—in foods and drugs, consumer products, lakes and streams, and even living cells. All biological cells depend on the properties of buffers, as does the essential function of the

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Properties of Buffer Solutions provides a culminating activity that ties together essential learning objectives relating to strong and weak acids, conjugate acid-base pairs, and more.

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Reasoning 1. The major buffer in blood is composed of the weak acid carbonic acid (H_2CO_3) and its conjugate base, bicarbonate ion (HCO_3^-).

Properties of Buffer Solutions: by Carissa Villanueva on ...

Explore the properties of biological buffers in

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this two-part kit. In Part 1, students examine the acid–base balance in a model carbonate “blood” buffer. The pH range over which buffers are effective is illustrated in Part 2, where students prepare two different phosphate buffers that reflect the physiological role of

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buffers within cells.

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Laboratory #19.

Introduction. One of

the most important

applications of acids

and bases in

chemistry and biology

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is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it.

pH Properties of
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In this laboratory, students investigate the properties of buffer solutions. The

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students make two ideal buffer solutions, one consisting of a weak acid and its conjugate base and the other, a weak base and its conjugate acid. The initial pH of each solution is determined. Strong acid and strong base are then added to each buffer in a series

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of steps, with the pH determined after each addition.

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Chemistry Big Idea 6

...

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Buffers Keep the
Balance—Properties of
Biological Buffers ...
Lab #16 - Properties
of Buffer Solutions.
The weak acid

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component HA reacts with any base added to the solution to give its conjugate base A^- . The conjugate base component A^- reacts with any acid added to the solution to form its acid partner HA. These reaction can be visualized as a cyclic process (see Figure 1 below).

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