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II QUANTITATIVE CORRELATION OF RATES AND EQUILIBRIA References: 1. Carey and Sundberg, Part A, pp. 196-209 (3rd ed.) (4th Ed. pp. 204-215) (5th Ed. pp. 335-344) 2. Carroll, p. 371 3. Miller, p.122 4. L&R pp. 143-158 5. Isaacs Chapter 4 6. Anslyn/Dougherty, pp 445-451 6. J. Hine, Structural Effects on Equilibria in Organic Chemistry (see p.1) 7.

CHAPTER 1 SUBSTITUENT EFFECTS ON ORGANIC RATES AND EQUILIBRIA

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$r = \frac{1}{a} \frac{dC_A}{dt} = \frac{1}{b} \frac{dC_B}{dt} = \dots$
from the expression above, it is clear that the usual convention is that reaction rate is taken as products formation rate. Rate is a function of the concentration of reactants and products, temperature, pressure and presence of a catalyst.

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libraries, and are recommended for private purchase by graduate students and professional organic chemists. Joseph F. Bunnett Brown University Providence, Rhode Island ...

6.7: Describing a Reaction - Equilibria, Rates, and Energy ...

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Objectives. After completing this section, you should be able to. write the equilibrium constant expression for a given reaction. assess, qualitatively, how far a reaction will proceed in a given direction, given the

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value of K_{eq} ; explain the difference between rate and equilibrium.

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