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Hybridization of ClF3: Hybridization of Cl in Chlorine ...

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What are the Two Types of Ions and ... - RS Aggarwal Solutions

The central atom Cl needs three unpaired electrons to bond with three F-atoms. ClF3 should consist of 3 bond-pairs and 2 lone-pairs. One 3s, three 3p and one of the 3d orbitals of Cl participate in the hybridization and five sp 3 d hybrid orbitals are formed. ClF3 Molecular Geometry And Bond Angles. ClF3 molecular geometry is said to be a T-shaped.

MCQ Questions for Class 11 Chemistry Chapter 8 Redox ...

For example, hydrochloric acid in the water. HCl undergoes dissociation reaction to produce H + ion and Cl – ion, as explained below. The concentration of the H+ ions is increased by forming hydronium ion. $\text{HCl (aq)} \rightleftharpoons \text{H}^+ \text{(aq)} + \text{Cl}^- \text{(aq)}$ $\text{HCl (aq)} + \text{H}_2\text{O (l)} \rightleftharpoons \text{H}_3\text{O}^+ \text{(aq)} + \text{Cl}^- \text{(aq)}$ Other examples of Arrhenius acids are listed below

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What are the Two Types of Ions and how are they Different Ion: An ion is a positively or negatively charged atom (or group of atoms). An ion is formed by the loss or gain of electrons by an atom, so it contains an unequal number of electrons and protons. Example: Sodium ion Na+, magnesium ion [...]

MCQ Questions for Class 10 Science Chapter 2 Acids, Bases ...

The oxidation number of Cl in Cl 2 O 7 is (a) + 7 (b) + 5 (c) + 3 (d) – 7. Answer. Answer: (a) + 7 Explanation: Cl show different oxidation state as -1 to +7 due to vacant d orbital. As oxygen is more electronegative than Cl. Oxygen size is small hence its more electronegative and show -2 oxidation states. Here Cl 2 O 7 then equation is: 2x ...

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