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In the method of initial rates that will be used in the experiment, you will measure the time it takes for the first small percentage of the reaction to occur. In our reacting system, this corresponds to the time it takes for the thiosulfate ion (S₂O₃²⁻) to be used up (see background section of experiment for details). For this to work, the amount of thiosulfate anion should be small relative ...

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The simplest acid-base reactions are those of a strong acid with a strong base. Table 4 shows data for the titration of a 25.0-mL sample of 0.100 M hydrochloric acid with 0.100 M sodium hydroxide. The values of the pH measured after successive additions of small amounts of NaOH are listed in the first column of this table, and are graphed in Figure 1, in a form that is called a titration curve.

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