

Specific Heat Worksheet 2 Answers

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Specific Heat Worksheet - Socorro Independent School ...
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Worksheet- Calculations involving Specific Heat
Name Answer Key Date 9/9/15 Chp 2-1: Specific Heat Worksheet
 $(m)(\Delta T)(C_{sp})=Q$ 1. Specific heat is the amount of energy that it takes to raise the temperature of 1 gram of a substance by 1 degree kelvin

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Specific Heat and Heat Capacity Worksheet

Specific Heat Worksheet #2 (Answer Key) 1. Brass is an alloy made from copper and zinc. A 0.66 kg sample of brass at 98.6 o C is dropped into 2.33 kg of water at 4.6 o C. If the equilibrium temperature is 7.0 o C, what is the specific heat capacity of brass.

Specific Heat Capacity Handout Answer Key

Specific Heat Calculations Worksheet. In a heat calculation problem, if the problem asks about melting/freezing you would multiply the mass times _____. heat of fusion. heat of vaporization. or specific heat. In a heat calculation problem, if the problem asks about a change in temperature, you would multiply the mass times _____ times the ...

Specific Heat Worksheet - Broadneck High School

Specific Heat and Heat Capacity Worksheet. 1 The temperature of 335 g of water changed from 24.5oC to 26.4oC. How much heat did this sample absorb? c for water = 4.18 J/goC 2. How much heat in kilojoules has to be removed from 225g of water to lower its temperature from 25.0oC to 10.0oC? 3.

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Activity—Specific Heat Capacity Handout Answer Key 2 4. To heat the hot chocolate to the optimal temperature of 57 ° C, how much energy is needed? $Q = mc \Delta T$ $Q = (50 \text{ g})(3.9 \text{ J/g } ^\circ \text{ C})(57 ^\circ \text{ C} - 40 ^\circ \text{ C})$ $Q = 3,315 \text{ J}$ Analysis Questions Answers will vary, depending on collected data. 1. Water has a specific heat of 4.18 J/g ° C.

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Specific Heat Worksheet Answer Key | Briefencounters

Worksheet- Calculations involving Specific Heat 1. For $q = mc \Delta T$: identify each variables by name & the units associated with it. $q =$ amount of heat (J) $m =$ mass (grams) $c =$ specific heat ($J/g \text{ } ^\circ C$) $T =$ change in temperature ($^\circ C$) 2. Heat is not the same as temperature, yet they are related. Explain how they differ from each other.

Specific Heat Capacity Worksheet (with answers) | Teaching ...

j ri phufxu\ lv khdwhg iurp & wr & dqg devruev mrxohv ri khdw lq wkh surfhvv & dofxodwh wkh vshflilf khdw fdsdflw\ ri phufxu\ :kdw lv wkh vshflilf khdw fdsdflw\ ri vloyhu phwdo li j ri wkh phwdo devruev - ri khdw

Specific Heat Worksheet-2Answer Key - Specific Heat ...

Specific Heat Practice Worksheet 1. An aluminum skillet weighing 1.58 kg is heated on a stove to $173 \text{ } ^\circ C$. Suppose the skillet is cooled to room temperature, $23.9 \text{ } ^\circ C$. How much heat energy (joules) must be removed to cause this cooling? The specific heat of aluminum is $0.901 \text{ J/(g} \cdot \text{ } ^\circ C)$. 2.

Specific Heat Chem Worksheet 16 1 Answer Key | Briefencounters

Object A specific heat is $2.45 \text{ J/g} \cdot \text{ } ^\circ C$ and object B specific heat is $0.82 \text{ J/g} \cdot \text{ } ^\circ C$. Which object will heat up faster if they have the same mass and equal amount of heat is applied? Explain why. 7. If $8.00 \text{ g NH}_4 \text{ NO}_3$ is dissolved in 1000 g of water, the water decreases in temperature from $21.00 \text{ } ^\circ C$ to $20.39 \text{ } ^\circ C$. Determine the molar heat of ...

Specific Heat Worksheet

Specific Heat Worksheet (m)(ΔT)(C_{sp})= Q ... Object A specific heat is $2.45 \text{ J/g} \cdot \text{ } ^\circ C$ and object B specific heat is $0.82 \text{ J/g} \cdot \text{ } ^\circ C$. Which object will heat up faster if they have the same mass and equal amount of heat is applied? Explain why. When a 120 g sample of aluminum (Al) absorbs 9612 J of energy, its temperature increases from $25 \text{ } ^\circ C$ to

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Specific Heat Worksheet 2 Answers

Answers are provided at the end of the worksheet without units. 1. A 15.75-g piece of iron sorbs 1086.75 joules of heat energy, and its temperature changes from 25.0 to 175.0°C. Calculate the specific heat capacity of iron. = 'C' Q 5) 2. How many joules of heat are needed to raise the temperature of 10.0 g of

Heat Transfer/ Specific Heat Problems Worksheet

Specific Heat Worksheet Answers from specific heat chem worksheet 16 1 answer key , source:mychaume.com. Informal together with feedback sessions help do away with splinters that may hamper the practice of achieving the vision. Adhere about what to edit to the instructions.

Heat Calculations Worksheet

Heat Transfer/ Specific Heat Problems Worksheet Solving For Heat

(q) 1. How many joules of heat are required to raise the temperature of 550 g of water from 12.0 °C to 18.0 °C? 2. How much heat is lost when a 64 g piece of copper cools from 375 °C, to 26 °C? (The specific heat of copper is 0.38452 J/g x °C). Place your answer in kJ. 3.

Specific Heat Wksht20130116145212867

View Specific Heat WS Answers from SCIENCE A at Southwestern Academy. Name Answer Key Date 9/9/15 Chp 2-1: Specific Heat Worksheet (m)(T)(Csp)=Q 1. Specific heat is the amount of energy that it takes

Specific Heat Worksheet Extra-1

17. The specific heat of ethanol is 2.46 J/g °C. Find the heat required to raise the temperature of 193 g of ethanol . from 19°C to 35°C. 18. When a 120 g sample of aluminum (Al) absorbs 9612 J of energy, its

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temperature increases from 25°C to 115°C. Find the specific heat of aluminum.

Chemistry*Temperature&SpecificHeat*Worksheet* Answer Key

Two page worksheet using Specific Heat Capacity. Questions start easy then become gradually harder. Answers included on separate sheet.

Also includes a spreadsheet to show how the calculations have been done.

Specific Heat Practice Worksheet

1) How much heat is required to raise 15.0 g of water from 20.0 °C to 50.0 °C? 2) A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron. 3) How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22 °C to 55 °C?

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