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What is stoichiometry? 1.

true 2. true 3. false 4.

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7. 3, 3, 96.00 8. 2, 2,

88.02 9. 4, 4, 72.08 10.

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methanol and oxygen gas 11.
carbon dioxide and water 12.
160.10 g 13. 160.10 g 14.
They are equal. 15. A mole
ratio is a ratio between ...

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quantitative relationships
between the amounts of
reactants used and the
amounts of products formed
by a chemical reaction is
called stoichiometry. _____
2. Stoichiometry is based on

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the law of conservation of mass. _____ 3.

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Stoichiometry Answer The
study of the quantitative
relationships between the
amounts of reactants used
and the amounts of products
formed by a chemical
reaction is called
stoichiometry. _____ 2.
Stoichiometry is based on
the law of conservation of
mass.

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Stoichiometric Calculations

In your textbook, read about
mole-to-mole conversion.

Read the following passage
and then solve the problems.

In the equation that follows
each problem, write in the
space provided the mole
ratio that can be used to
solve the problem.

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