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One mole of an ideal gas is heated at constant pressure from 0°C to 200°C. (a) Calculate work done. (b) If the gas were expanded isothermally & reversibly at 0°C from 1 atm to some other pressure  $P_t$ , what must be the final pressure if the maximum ...

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The third law of thermodynamics states that: "The entropy of all the perfect crystalline solids is zeros at absolute zero temperature". The third law of thermodynamics is also referred to as Nernst law. It provides the basis for the calculation of absolute entropies of the substances. Mathematically,  $\lim_{T \rightarrow 0} \dots$

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Chapter 17. Work, Heat, and the First Law of Thermodynamics

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Entropy 22.7 Entropy Changes in Irreversible Processes Scale 22.8 Entropy on a Microscopic Heat Engines, Entropy, and the Second Law of Thermodynamics

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The First Law of Thermodynamics Work and heat are two ways of transferring energy between a system and the environment, causing the system's energy to change. If the system as a whole is at rest, so that the bulk mechanical energy due to translational or rotational motion is zero, then the

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