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Properties of water - Wikipedia

An acid is a molecule or ion capable of donating a proton (hydrogen ion H +) (a Brønsted–Lowry acid), or, alternatively, capable of forming a covalent bond with an electron pair (a Lewis acid).. The first category of acids are the proton donors, or Brønsted–Lowry acids.In the special case of aqueous solutions, proton donors form the hydronium ion H 3 O + and are known as Arrhenius acids.

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Water has a very high specific heat capacity of 4.1814 J/(g·K) at 25 °C – the second highest among all the heteroatomic species (after ammonia), as well as a high heat of vaporization (40.65 kJ/mol or 2257 kJ/kg at the normal boiling point), both of which are a result of the extensive hydrogen bonding between its molecules. These two unusual properties allow water to moderate Earth's ...